**TYPES OF CHEMICAL EQUATIONS (IDENTIFICATION)**

**State whether each of the following equations represents a synthesis (s) , decomposition(d), single replacement (sr), double replacement (dr), or combustion reaction (c).**

\_\_\_\_\_\_\_\_\_\_\_\_\_ 1. CO2 → C + O2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 2. NaCl + AgNO3 → NaNO3 + AgCl

\_\_\_\_\_\_\_\_\_\_\_\_\_ 3. S + Cl2 → SCl2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 4. BaCl2 + 2NaOH -→ 2NaCl + Ba(OH)2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 5. 2C2H2 + 5O2  → 4CO2 + 2 H2O

\_\_\_\_\_\_\_\_\_\_\_\_\_ 6. Zn + CuSO4 → ZnSO4  + Cu

\_\_\_\_\_\_\_\_\_\_\_\_\_ 7. CH4 → C + 2H2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 8. Pb(NO3)2 + Mg → Pb + Mg(NO3)2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 9. Mg + 2 HCl → MgCl2 + H2

\_\_\_\_\_\_\_\_\_\_\_\_ 10. H2SO4 → H2 + S + 2O2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 11. 2O2 + N2 → N2O4

\_\_\_\_\_\_\_\_\_\_\_\_\_ 12. 3CaBr2 + 2Na3P → Ca3P2 + 6NaBr

\_\_\_\_\_\_\_\_\_\_\_\_\_ 13. 2KI + Br2 → 2KBr + I2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 14. C6H12O6 → 6C + 6H2O

\_\_\_\_\_\_\_\_\_\_\_\_\_ 15. 2NaF → 2Na + F2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 16. Si + O2 → SiO2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 17. 2NaI + Pb(NO3)2 →2NaNO3 + PbI2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 18. H2 + CO + O2 → H2CO3

\_\_\_\_\_\_\_\_\_\_\_\_\_ 19. C3H8 + 5O2 → 3CO2 + 4H2O

\_\_\_\_\_\_\_\_\_\_\_\_\_ 20. Li3PO4 → 3Li + P +2O2

\_\_\_\_\_\_\_\_\_\_\_\_\_ 21. CS2 + 2F2 → CF4 + 2S

\_\_\_\_\_\_\_\_\_\_\_\_\_ 22. NaI + Cs → CsI + Na